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Reactions 2 4.

Surface area of

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reactants - exposed surfaces affect speed. 0
Except for substances in the gaseous state or solution, reactions occur at the boundary, or interface, between two phases.

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KINETICS

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This is a

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presentation and
11-page guided
notes packet
used to explain
some of the
fundamental
concepts of

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kinetics. It

covers various

reaction rate

topics

including: zero,

first, and

second order

rates,

calculating

using the method

of initial

rates, half

lives, colli.

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$2\text{O}_2(\text{g}) \rightarrow 2\text{N}_2\text{O}(\text{g}) +$

$\text{O}_2(\text{g})$ rate =

$k[\text{N}_2\text{O}]^2$. For an

initial

concentration of

N_2O of 0.50 M ,

calculate the

concentration of

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N. 20 remaining after 2.0 min if $k = 3.4 \times 10^{-3} \text{ s}^{-1}$. 15. A reaction has a rate constant of 0.00300 s^{-1} .

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Practice

Problems and

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Dr. Mandes The following questions represent potential types of quiz questions.

Please answer each question completely and thoroughly. The solutions will be posted on-line on Monday.

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5. Please do #18
in chapter 12 of
your text. a.

KINETICS

Practice

Problems and
Solutions

A.P. Chemistry

Practice Test:

Ch. 12, Kinetics

MULTIPLE CHOICE.

Choose the one
alternative that

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best completes
the statement or
answers the
question. 1)

Consider the
following
reaction: $3A \rightarrow 2B$
The average
rate of
appearance of B
is given by
 $D[B]/Dt$.

Comparing the
rate of

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Answers

appearance of B
and the rate of
disappearance of
A, we get

$$D[B]/Dt = \frac{\quad}{x (-D[A]/Dt)}.$$

A.P. Chemistry
Practice Test:
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MULTIPLE ...
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Kinetics 1. The rate equation for a chemical reaction is determined by

- (A) theoretical calculations.
- (B) measuring reaction rate as a function of concentration of reacting species.
- (C) determining the

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equilibrium constant for the reaction. (D) measuring

reaction rates as a function of temperature 2.

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Unit 3: Chemical

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Questions page 2
Mechanism 3 is
correct. The
rate law shows
that the slow
reaction must
involve one Y,
consistent with
mechanism 3.

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Mechanisms 1 and 2 would involve both [X] and [Y] in the rate law, not consistent with the rate law. 1987 a) three points; one each for form of rate law, HgCl₂ exponent, C₂O₄²⁻ exponent

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Grade. In this

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instructional

activity,

students find

rate laws of

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reactions, they determine rate constants, they find rate-determining steps of reactions, they find half-life's of reactions and they determine the overall order of reactions.

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Equal numbers of
moles of F. 2

(g) and ClO. 2

(g) are drawn

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into a vacuum
where the
following
process takes
place. F Lecture
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study of rates

of chemical

reactions and

the mechanisms

by which they

occur FACTORS

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WHICH AFFECT
REACTION RATES

1. Nature of the
Reactants – more
active
substances react
faster (ie ...

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covered during
the first year
course and
reviewed over a
short time span
during the AP
course. (Links
for these units
can be found
under the
"CHEMISTRY" tab
above and on the
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overview.)

Units 1-5 Pre-AP
Chemistry Review
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...

6 Kinetics and
Equilibrium
Chemical
kinetics can be
divided into two
parts. The
first, at the

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macroscopic level, is the study of rates of reactions: what the rate of reaction means; how to determine a rate by experiment; and how factors, such as the concentrations of reactants and temperature,

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