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Formulas. STUDY. PLAY. Monatomic
Ion. Consists of a single atom with a
positive or negative charge resulting from

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the loss of gain of one or more valence electrons and behaves as a unit, yet it still has a charge. (+ or -). Determined by the amount of electrons lost or gained.

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29 terms. hales913. Chapter 9: Chemical

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Section 9.4 – Naming and Writing

Formulas for Acids and Bases. An acid is a compound that produces H^+ ions when it dissolves in water. The formula for an acid normally starts with and H. When naming acids, you should first determine whether or not the acid contains oxygen.

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and Formulas. Acids. base. Cation. Ionic
compounds. Compounds that contain one
or more hydrogen atoms and produce....
An ionic compound that produces
hydroxide ions when dissolved.... Any

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atom or group of atoms that has a positive charge. Compounds composed of cations and anions.

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83 SECTION 9.3 NAMING

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AND WRITING FORMULAS FOR
MOLECULAR COMPOUNDS (pages
268 – 270) This section explains the rules
for naming and writing formulas for
binary molecular compounds. Naming
Binary Molecular Compounds (pages
268 – 269) 1. Circle the letter of the type(s)
of elements that form binary molecular

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Answers

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SECTION 9.2 NAMING AND
WRITING FORMULAS FOR IONIC
COMPOUNDS I. Write the formulas for
these binary ionic compounds. c.

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potassium iodide 2. write the formulas for
the compounds formed c. b. f. 3. Name
the following binary ionic compounds.

sodium sulfide g. CuCl : h. SnCl . i. a. MnO_2
c. d. SrBr_2 221 Chapter 9 Chemical
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9.1 Naming Ions 9.2 Naming and Writing

Formulas for Ionic Compounds 9.3

Naming and Writing Formulas for

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The Laws Governing How Compounds

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Formulas ” ... in parenthesis after the
name of the metal (Table 9.2, p.255)
Predicting Ionic Charges Some of the post-
transition elements also have more than

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one possible oxidation state. Tin (II) = Sn^{2+}
Lead (II) = Pb^{2+} Tin (IV) = Sn^{4+}
Lead (IV) = Pb^{4+}

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CHAPTER 9 Assessment 42. a. 2 – b. 1+
c. 1 – d. 3+ 43. a. 2+ b. 2+ c. 3+ d. 1+ 44.
a. barium ion b. iodide ion c. silver ion d.
mercury(II) ion 45. cyanide, CN^- and
hydroxide, OH^- 46. a. hydroxide ion b.
lead(IV) ion c. sulfate ion d. oxide ion 47.

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zero; A compound is electrically neutral.
48. The symbols for the cation and

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Chapter 9 Practice Problems: Chemical
Names and Formulas. Section 9.1:

Naming Ions. 1. What is the charge on the
ion typically formed by each element? a.

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oxygen c. sodium e. nickel, two electrons
lost . b. iodine d. aluminum f. magnesium.

2. How many electrons does the neutral
atom gain or lose when each ion forms? a.
Cr³⁺ c. Li⁺ e. Cl⁻ b. P³⁻ ...

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Formulas, the masses of the elements are always in the same proportions any atom or group of atoms that has a negative charge ... place cation name first followed by the anion name: naming polyatomic ...

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